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Place the ions in their proper order: cation and then anion. Name the cation. Metals that form only one cation. As noted in Section 3.3, these metals are usually in groups 1–3, 12, and 13. The name of the cation of a metal that forms only one cation is the same as the name of the metal (with the word ion added if the cation is by itself). For example, Na⁺ is the sodium ion, Ca²⁺ is the ...

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Chapter 10 Nuclear Chemistry Section 10.3 Artificial ...

Section 10.1 Measuring Matter Section 10.2 Mass and the Mole Section 10.3 Moles of Compounds Section 10.4 Empirical and Molecular Formulas Section 10.5 Formulas of Hydrates Exit CHAPTER Table Of Contents 10 Click a hyperlink to view the corresponding slides. • Explain how a mole is

Chemistry: Matter and Change

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and

nonconducting

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SECTION 1 3. Gas particles are in continuous, rapid, random motion. They therefore possess kinetic energy, which is energy of motion.

CHAPTER 10 States Matter

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SECTION 3 Date CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (b) divided by the mass of 1 mol. nRT (c) multiplied by 22.4 L. (d) divided by 22.4 L. For the expression $V =$ (a) increasing P

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