

Chemistry Chapter 12 Solution Manual Stoichiometry

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CHAPTER 12 SOLUTIONS MANUAL States of Matter Section 12.1 Gases pages 402-410 Practice Problems page 405 1. Calculate the ratio of effusion rates for nitrogen (N₂) and neon (Ne). Rate N / Rate Ne = 0.849 Rate N / Rate Ne = $\sqrt{\frac{\text{molar mass Ne}}{\text{molar mass N}_2}}$ = $\sqrt{\frac{20.18}{28.02}}$ = 0.849 2. Calculate the ratio of diffusion rates for carbon monoxide and carbon dioxide. 1.25 Rate CO / Rate CO₂ = $\sqrt{\frac{\text{molar mass CO}_2}{\text{molar mass CO}}}$

States of Matter

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4.2 For the reaction, $2A + B \longrightarrow A_2B$, the rate = $k [A][B]^2$ with $k = 2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$. Calculate the initial rate of the reaction when $[A] = 0.1 \text{ mol L}^{-1}$, $[B] = 0.2 \text{ mol L}^{-1}$. Calculate the rate of reaction after $[A]$ is reduced to 0.06 mol L^{-1} . Sol. Initial rate of reaction = $k [A] [B]^2 = (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ s}^{-1}) (0.1 \text{ mol L}^{-1}) (0.2 \text{ mol L}^{-1})^2 = 8 \times 10^{-9} \text{ mol L}^{-1} \text{ s}^{-1}$.

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